Breaking the Table Code HW

Read and outline (Cornell Style) **Section 6.3** in your chemistry textbook. Then answer the following assessment questions:

1. On your periodic table (in the front of your notebook), label and color the following types of elements:
    a. Alkali metals
    b. Alkaline earth metals
    c. Transition metals
    d. Other metals
    e. Metalloids
    f. Halogen nonmetals
    g. Noble gas nonmetals
    h. Other nonmetals
    i. Lanthanide metals
    j. Actinide metals
2. Sketch a simplified periodic table and use arrows and labels to compare period and group trends in atomic radii (size), ionization energy, electronegativity, atomic mass, atomic number, and # of valence electrons.
3. Which has the largest atomic radius: nitrogen (N), antimony (Sb), or arsenic (As)? The smallest?
4. For each of the following properties, indicated whether fluorine (F) or bromine (Br) has a larger value.
    a. Electronegativity
    b. Ionic radius
    c. Atomic radius
    d. Ionization energy